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determination of rate laws. First-order reaction (with calculus) Plotting data for a first-order reaction.

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statement or answers the question. 1)  
Consider the following reaction:  $3A \rightarrow 2B$   
The average rate of appearance of B is given by  $\frac{1}{2} \frac{d[B]}{dt}$ . Comparing the rate of appearance of B and the rate of. [https://pobchemteam.weebly.com/uploads/2/5/7/7/25772138/ch\\_12\\_prac\\_test\\_kinetics1.pdf](https://pobchemteam.weebly.com/uploads/2/5/7/7/25772138/ch_12_prac_test_kinetics1.pdf) read more.

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curry.zerohate.me Chemical kinetics - Chemistry test 1) Chemical kinetics is the branch of chemistry which deals with the study of: a) Speed or rate of chemical reaction. b) The factors affecting the rates of the reaction. Chemical kinetics - Chemistry test

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Answers: 1. (a) When we compare the results of experiments 2 and 3, we see that when  $[B]$  doubles, the rate doubles, so the reaction is first order with respect to B. Experiments 2 and 3 prove that the rate must double when  $[B]$  doubles. Knowing this, we can see that when the rate doubles from experiment 1 to 2, it must be because of B, and the

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## **AP Chemistry Self-Test Worksheet Kinetics**

Answer Key Testname:

CH\_12\_PRAC\_TEST\_KINETICS.TST

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) C ID: chem9b 14.1-1 2) B ID: chem9b

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14.1-2 3) A ID: chem9b 14.1-3 4) E ID:  
chem9b 14.1-9 5) B ID: chem9b 14.1-19  
6) B ID: chem9b 14.1-24 7) E ID: chem9b  
14.1-26 8) C ID: chem9b 14.1-28 9) D ID:  
chem9b 14.1-37 10) B

## **A.P. Chemistry Practice Test: Ch. 12, Kinetics MULTIPLE ...**

Chemistry 12 Kinetics Practice Test # 2

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1. Which of the following units could be used to express the reaction rate? A. mL/s B. mL/g C. g/mL D. mL/mol

2. Consider the reaction:  $\text{Zn (s)} + 2\text{HCl (aq)} \rightarrow \text{ZnCl}_2\text{(aq)} + \text{H}_2\text{(g)}$  The rate of production of  $\text{ZnCl}_2$ , can be increased by A. decreasing the  $[\text{HCl}]$ . B. increasing the temperature

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Energy, Transition States... P10

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## **4.1 Kinetics - A-Level Chemistry**

4.2 For the reaction,  $2A + B \longrightarrow A_2B$ , the rate =  $k [A][B]^2$  with  $k = 2.0 \times$

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$10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$ . Calculate the initial rate of the reaction when  $[A] = 0.1 \text{ mol L}^{-1}$ ,  $[B] = 0.2 \text{ mol L}^{-1}$ . Calculate the rate of reaction after  $[A]$  is reduced to  $0.06 \text{ mol L}^{-1}$ . Sol. Initial rate of reaction =  $k [A] [B]^2 = (20 \times 10^{-6} \text{ mol}^{-2} \text{ s}^{-1}) (0.1 \text{ mol L}^{-1}) (0.2 \text{ mol L}^{-1})^2 = 8 \times 10^{-9} \text{ mol L}^{-1} \text{ s}^{-1}$ .

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Chemical kinetics - Chemistry test 1)

Chemical kinetics is the branch of chemistry which deals with the study of:  
a) Speed or rate of chemical reaction. b)

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The factors affecting the rates of the reaction.

## **Chemical kinetics - Chemistry test**

Chemical kinetics - Chemistry test - Part 2 11) The collision frequency is directly proportional to the: a) Absolute temperate. b) Square root of the absolute temperature. c) Square of

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absolute temperature. d) None of these.  
ANSWER : Square root of the absolute temperature.

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