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Chapter 19: Reaction Rates and Equilibrium  
19.  $2\text{CH}_4(\text{g}) \rightleftharpoons \text{C}_2\text{H}_2(\text{g}) + 3\text{H}_2(\text{g})$   
While studying this reaction mixture, a chemist analysed a 4.

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a state of balance in which the rates of the forward and reverse reactions are equal; no net change in the amount of reactants and products occurs in the chemical system (18.2) equilibrium position the relative concentrations of reactants and products of a reaction that has reached equilibrium; indicates whether the reactants or products are ...

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Rates of Reactions and Equilibrium. The rate of reaction and the factors affecting it is a key topic in the GCSE chemistry specifications. You need to understand how these different factors such as pressure, concentration, temperature

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and the presence of a catalyst impact on the equilibrium of a reversible reaction.

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about things like the field that studies ...

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The equilibrium constant,  $K_{eq}$  for the reaction below is 85.0 at 460°C. If a mixture is prepared where the initial concentration of sulfur dioxide is 1.00 mol/L and nitrogen dioxide is 2.00 mol/L

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calculate the equilibrium concentration of nitrogen monoxide and nitrogen dioxide.

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This unit includes resources, such as slideshows, worksheets, and labs, on Reaction Rates and Equilibrium for high

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the four factors that affect reaction rates. temperature. ... if a stress is applied to a system in equilibrium, the reaction will shift in the direction to

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AND THE EQUILIBRIUM CONSTANT Many

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chemical reactions do not go to completion but instead attain a state of chemical equilibrium. Chemical equilibrium: A state in which the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant.

## **Chapter 14. CHEMICAL EQUILIBRIUM**



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In this kinetics and equilibrium worksheet, students use the diagram to answer the questions as it relates to endothermic or exothermic reactions. Students indicate the result in the reaction rate as speeding up or slowing down.

## **Kinetics and Reaction Rates Lesson**

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Reaction Rates in Analysis: Test Strips for Urinalysis. Physicians often use disposable test strips to measure the amounts of various substances in a patient's urine ( ). These test strips contain various chemical reagents, embedded in small pads at various locations along the strip, which undergo

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changes in color upon exposure to sufficient concentrations of specific substances.

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